

# ТВЕРСКИЯ ГУБЕРНСКИ ВЪДМОСТИ



ПЕРВОЕ ВЫХОДИТ ВЪ СРЕДУ И ВЪ СУББОТУ

Въ Твери	№ 10	1907
Въ Ярославлѣ	№ 10	1907
Въ Москвѣ	№ 10	1907

Въ Калужѣ	№ 10	1907
Въ Рязанѣ	№ 10	1907
Въ Смоленскѣ	№ 10	1907

№ 10      Четвертъ 1907 г.      № 10

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### О выдѣлѣ земель въ Тверской губернии

Въ Тверской губернии выдѣлены земли, принадлежащія на праве собственности Государственному банку, въ слѣдующихъ уѣздахъ:

1. **Тверской уѣздъ**

2. **Селифериинскій уѣздъ**

3. **Спасскій уѣздъ**

4. **Торжокскій уѣздъ**

5. **Угличскій уѣздъ**

6. **Ярославскій уѣздъ**

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### О выдѣлѣ земель въ Ярославской губернии

Въ Ярославской губернии выдѣлены земли, принадлежащія на праве собственности Государственному банку, въ слѣдующихъ уѣздахъ:

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1. **Introduction**  
 This report discusses the results of the experiment conducted on the effect of temperature on the rate of reaction between hydrogen peroxide and potassium iodide. The reaction is exothermic and produces iodine and water. The rate of reaction was measured by the time taken for a fixed amount of iodine to be produced, which was then measured by the change in color of a starch solution.

**Method:** A series of test tubes were prepared containing different volumes of hydrogen peroxide, potassium iodide, and starch solution. The temperature of the reaction mixture was varied by adding different volumes of water. The time taken for the reaction to occur was measured by the time taken for the solution to turn blue-black.

**Results:** The results show that the rate of reaction increases as the temperature increases. This is because the molecules have more kinetic energy and are able to overcome the activation energy barrier more easily.

**Conclusion:** The experiment demonstrates that the rate of reaction between hydrogen peroxide and potassium iodide is directly proportional to the temperature of the reaction mixture.

2. **Method**  
 The experiment was carried out using a series of test tubes. A fixed volume of hydrogen peroxide was added to a fixed volume of potassium iodide solution. The temperature of the reaction mixture was varied by adding different volumes of water. The time taken for the reaction to occur was measured by the time taken for the solution to turn blue-black.

**Results:** The results show that the rate of reaction increases as the temperature increases. This is because the molecules have more kinetic energy and are able to overcome the activation energy barrier more easily.

**Conclusion:** The experiment demonstrates that the rate of reaction between hydrogen peroxide and potassium iodide is directly proportional to the temperature of the reaction mixture.

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